Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Date: \_\_\_\_\_\_\_\_\_\_\_\_\_\_

.3 moles KI = \_\_\_\_\_\_\_\_ grams

7.8 x 1025 atoms of cobalt = \_\_\_\_\_\_\_\_ grams

4.5 moles xenon = \_\_\_\_\_\_\_\_ grams

7.6 moles magnesium iodide = \_\_\_\_\_\_\_\_ grams

0.679 moles barium hydroxide = \_\_\_\_\_\_\_\_ grams

67.8 grams lead (II) oxide = \_\_\_\_\_\_\_\_ moles

0.75 moles potassium chloride = \_\_\_\_\_\_\_\_ grams

0.25 moles sodium iodide = \_\_\_\_\_\_\_\_ molecules

47.3 grams lead (IV) oxide = \_\_\_\_\_\_\_\_ atoms

6.8 moles carbon dioxide = \_\_\_\_\_\_\_\_ grams

675 grams osmium = \_\_\_\_\_\_\_\_ atoms

0.54 moles titanium = \_\_\_\_\_\_\_\_ grams

Challenge - Chlorophyll (the fun green stuff in plants) has the chemical formula C55H72MgN4O5.

Determine the formula mass of chlorophyll.

If you have 670 grams of chlorophyll, how many moles do you have?

How many molecules is in that 670 grams of chlorophyll?

How many grams of carbon dioxide are there in 5.67 moles?

How many moles of magnesium chloride are in 56.75 grams?

If I have 45.87 moles of Lead how many atoms do I have?

Little Jimmie went to the store and bought 2.34 x 1025 atoms of gold. How many moles of gold did little Jimmie buy?

How many grams of gold did Little Jimmie buy?

Determine the percent composition of sulfur in sulfuric acid. (H2SO4)

Calculate the Percent Composition of the following compounds.

a.       SO3

b. NaClO3

c. Ca(NO3)2

d. Ammonium Sulfate

The % mass carbon in CO2

A 34.80-g sample contains 3.83 g iron and 10.97 g bromine. Find the % of Fe and % of Br.

The % mass oxygen in SO3

Find the % mass nitrogen in NH3

Find the % mass sulfur in SO2

The % mass sodium in NaOH

Find the % mass sulfur in SO3